

FERRIALLANITE-(Ce), $\text{CaCeFe}^{3+}\text{AlFe}^{2+}(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, A NEW MEMBER OF THE EPIDOTE GROUP: DESCRIPTION, X-RAY AND MÖSSBAUER STUDY

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ABSTRACT

Ferriallanite-(Ce), ideally $\text{CaCeFe}^{3+}\text{AlFe}^{2+}(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, is the analogue of allanite-(Ce) with Fe^{3+} dominant in the octahedral site $M(1)$. It is of metasomatic origin and occurs in an alkaline granitic pegmatite of Mount Ulyn Khuren in the Altai Range (about $48^\circ 32' \text{N}$, $92^\circ 55' \text{E}$), Mongolian People's Republic. Ferriallanite-(Ce) is black, opaque or translucent, with a brown streak; the luster is resinous. The Mohs hardness is 6; it is brittle, without observed cleavage or parting. The fracture is conchoidal to uneven. Its density is 4.22 g/cm^3 . It is biaxial (-), $\alpha 1.825(2)$, $\beta 1.855(5)$, $\gamma 1.880(5)$, $2V_{\text{calc}} 83^\circ$. The strongest five peaks in the X-ray powder-diffraction pattern [$d(\text{Å})(hkl)$] are: $2.93(65)(113)$, $2.72(80)(120,013)$, $2.63(60)(\bar{3}11)$, $2.18(100)(122, \bar{1}23, \bar{4}01)$, $2.14(80)(221, \bar{4}03, \bar{2}23, 014)$. The crystal structure has been refined to $R = 0.026$ using 2296 observed reflections collected with a four-circle X-ray diffractometer [$a 8.962(2)$, $b 5.836(2)$, $c 10.182(2) \text{ Å}$, $\beta 115.02(1)^\circ$, space group $P2_1/m$]. The position of the hydrogen atom has been found. On the basis of the chemical data, structure refinement, Mössbauer spectroscopy and charge-distribution calculations, the cations have been assigned to the independent crystallographic sites and the following crystal chemical formula has been obtained: $(\text{Ca}_{0.97}\text{Ce}_{0.03})(\text{Ce}_{0.89}\text{Ca}_{0.11})(\text{Fe}^{3+}_{0.80}\text{Ti}^{4+}_{0.14}\text{Al}_{0.06})(\text{Al}_{0.56}\text{Fe}^{3+}_{0.44})(\text{Fe}^{2+}_{0.93}\text{Mn}_{0.07})(\text{Si}_{0.94}\text{Al}_{0.06}\text{O}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$. In the same sample, a single-point electron-microprobe analysis reveals an ideal composition $\text{CaCeFe}^{3+}_2\text{Fe}^{2+}(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, but not enough material is available for a full characterization of this potentially new epidote-group mineral.

Keywords: ferriallanite-(Ce), new mineral species, crystal structure, Mössbauer spectroscopy, allanite subgroup, epidote group, alkaline granitic pegmatite, Altai Range, Mongolia.

SOMMAIRE

La ferriallanite-(Ce), de composition idéale $\text{CaCeFe}^{3+}\text{AlFe}^{2+}(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, est un analogue de l'allanite-(Ce) à dominance de Fe^{3+} dans le site octaédrique $M(1)$. Cette nouvelle espèce minérale, d'origine métasomatique, provient d'une pegmatite granitique alcaline au mont Ulyn Khuren, chaîne de montagnes Altai (environ $48^\circ 32' \text{N}$, $92^\circ 55' \text{E}$), en République Populaire de Mongolie. Les cristaux de ferriallanite-(Ce) sont noirs, opaques ou translucides, avec une rayure brune; l'éclat est résineux. La dureté de Mohs est 6; c'est un minéral cassant, sans clivage ou plan de séparation observé. La fracture est conchoïdale à inégale. Sa densité est 4.22 g/cm^3 . Elle est biaxe (-), $\alpha 1.825(2)$, $\beta 1.855(5)$, $\gamma 1.880(5)$, $2V_{\text{calc}} 83^\circ$. Les cinq raies les plus intenses du spectre de diffraction (méthode des poudres) [$d(\text{Å})(hkl)$] sont: $2.93(65)(113)$, $2.72(80)(120,013)$, $2.63(60)(\bar{3}11)$, $2.18(100)(122, \bar{1}23, \bar{4}01)$ et $2.14(80)(221, \bar{4}03, \bar{2}23, 014)$. Nous en avons affiné la structure cristalline jusqu'à un résidu R de 0.026 en utilisant 2296 réflexions observées prélevées avec un diffractomètre à quatre cercles [$a 8.962(2)$, $b 5.836(2)$, $c 10.182(2) \text{ Å}$, $\beta 115.02(1)^\circ$, groupe spatial $P2_1/m$]. Nous précisons la position de l'atome d'hydrogène. D'après les données chimiques,

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l'affinement de la structure, la spectroscopie de Mössbauer et les calculs de la distribution des charges, nous avons pu attribuer les cations aux sites cristallographiques indépendants, et nous obtenons la formule cristallichimique suivante: $(\text{Ca}_{0.97}\text{Ce}_{0.03})(\text{Ce}_{0.89}\text{Ca}_{0.11})(\text{Fe}^{3+}_{0.80}\text{Ti}^{4+}_{0.14}\text{Al}_{0.06})(\text{Al}_{0.56}\text{Fe}^{3+}_{0.44})(\text{Fe}^{2+}_{0.93}\text{Mn}_{0.07})(\text{Si}_{0.94}\text{Al}_{0.06}\text{O}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$. Dans le même échantillon, une analyse ponctuelle indique une composition idéale $\text{CaCeFe}^{3+}_2\text{Fe}^{2+}(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, mais dans un volume insuffisant pour permettre une caractérisation complète de ce qui semble être un nouveau membre du groupe de l'épidote.

(Traduit par la Rédaction)

Mots-clés: ferriallanite-(Ce), nouvelle espèce minérale, structure cristalline, spectroscopie de Mössbauer, sous-groupe de l'allanite, groupe de l'épidote, pegmatite granitique alcaline, chaîne de montagnes Altai, Mongolie.

INTRODUCTION

Minerals belonging to the epidote group are represented by the general formula $A_2M_3(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, where *A* stands for Ca, REE, Sr, Mn and Pb, and *M* represents Al, Fe²⁺, Fe³⁺, Mn, Mg and V; some F can substitute for OH. All known members are monoclinic *P2₁/m* except for zoisite, which is orthorhombic *Pnma* and in polytypic relationship with clinozoisite (Ray *et al.* 1986). In the following minerals of the epidote group, the rare-earth elements (REE) are dominant in one of the two independent *A* sites, which is usually labeled *A*(2): allanite-(Ce) (Hintze 1897), androsite-(La) (Bonazzi *et al.* 1996), dissakisite-(Ce) (Grew *et al.* 1991), dollaseite-(Ce) (Peacor & Dunn 1988) and khristovite-(Ce) (Pautov *et al.* 1993); allanite-(La) and allanite-(Y) have been introduced by Levinson (1966) on the basis of data reported in the literature. A different pattern of cation occupancy at one of the three independent *M* sites (two in zoisite) generates different mineral species.

Allanite-(Ce), ideally $\text{CaCeFe}^{2+}\text{Al}_2(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, is characterized by the replacement of $\text{Ca}^{2+} + \text{Al}^{3+}$ by $\text{REE}^{3+} + \text{Fe}^{2+}$ and, with the minimum content of Fe²⁺, becomes $\text{Ca}(\text{Ce}_{0.51}\text{Ca}_{0.49})(\text{Al}_{2.49}\text{Fe}^{2+}_{0.51})(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$; some Fe²⁺ can be replaced by other divalent ions (mainly Mn²⁺ and Mg²⁺). The dominance of Fe³⁺ in the octahedral site *M*(1) leads to the new species ferriallanite-(Ce), ideally $\text{CaCeFe}^{3+}\text{AlFe}^{2+}(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, approved by IMA (2000–041) and described in this paper. The type material is labelled TsH-9026 and deposited in the Fersman Mineralogical Museum, Moscow, Russia.

OCCURRENCE

Ferriallanite-(Ce) has been found in the “Nepri-metnyi” (= Indiscernible in English) alkaline granitic pegmatite occurring in the northern slope of Mount Ulyn Khuren (approximately 48°32'N, 92°55'E), which belongs to the Khaldzan Buragtag peralkaline granite massif, exposed 55 km north-northeast of Kobdo city, Mongolian Altai range, Kobdo aimak, Mongolian People's Republic. The new mineral is of metasomatic origin and represents a major product of the pseudomorphism of REE-rich minerals of the eudialyte group. It

occurs as an aggregate of subhedral grains from 0.3 × 0.5 to 1 × 2 mm within masses of zircon, quartz and kainosite-(Y) and is associated with aegirine, β-fergusonite-(Y), Y-rich and normal ilvaite, hingganite-(Ce), Nd-rich allanite-(Ce), magnetite, fayalite, fluorite.

EXPERIMENTAL

Physical properties

Ferriallanite-(Ce) is black and opaque, translucent (red or orange-red inner reflections), with a brown streak; the luster is resinous; fluorescence was not observed. It is brittle and shows a hardness $\text{VHN}_{100\text{g}}$ of 1250 kg/mm², corresponding to 6 on the Mohs scale. No cleavage or parting has been observed; the fracture is from conchoidal to uneven. By the microvolumetric method, a density of 4.22 g/cm³ was measured; the calculated density is 4.16 and 4.21 g/cm³ for the composition of the crystal used for structure refinement and that corresponding to column 2 in Table 1, respectively.

At a wavelength of 589 nm, ferriallanite-(Ce) is biaxial (-), with α 1.825(2), β 1.855(5) and γ 1.880(5); $2V$ could not be measured because of the black color; $2V_{\text{calc}}$ is equal to 83°. The dispersion is strong and shows $r < v$; pleochroism is observed: $Z > Y > X$, with *Z* dark red-brown, *Y* brown and *X* greenish gray. The compatibility index, 0.007, is superior.

The infrared absorption spectrum of ferriallanite-(Ce) shows absorption bands at the following wavelengths: 425, 465, 487, 529, 551, 569, 615, 681, 830, 864, 888, 919, 1033, and 3150 cm⁻¹.

Mössbauer spectra and chemical composition

Substantially identical Mössbauer spectra (Fig. 1) have been collected at room temperature from the type sample (TsH-9026) and a second one (TsH-9031) with an absorber thickness of 5 mg Fe/cm². The spectra were collected in transmission mode on a constant-acceleration spectrometer using a nominal 1.85 GBq ⁵⁷Co source with a velocity scale calibrated with respect to α-iron. They could be unambiguously fitted (program NORMOS written by R.A. Brand and distributed by Wissenschaftliche Elektronik GmbH, Germany) to two Lorentzian quadrupole doublets for Fe³⁺ and one

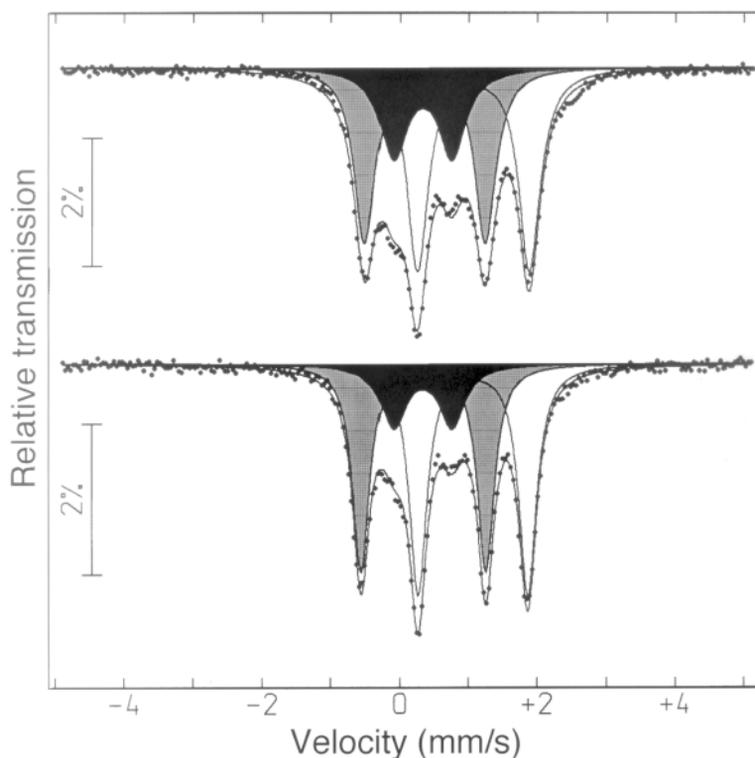


FIG. 1. Room-temperature Mössbauer spectra for two samples, TsH-9026 (top) and TsH-9031 (bottom) of ferriallanite-(Ce). The doublet corresponding to Fe^{2+} is unshaded, and the two doublets corresponding to Fe^{3+} are shaded black and gray.

Lorentzian quadrupole doublet for Fe^{2+} . Conventional constraints (component widths and areas constrained to be equal) were applied. Effects due to thickness and differences in the recoil-free fraction were estimated and found to be within the quoted experimental uncertainty. Average values of the center shifts [1.068(5) mm/s for Fe^{2+} ; 0.359(5) and 0.334(5) mm/s for Fe^{3+} , relative to α -iron], the quadrupole splittings [1.62(1) mm/s for Fe^{2+} ; 1.78(1) and 0.85(1) mm/s for Fe^{3+}], and the number of the fitted doublets show that: (i) Fe^{2+} occupies only one octahedral site, which is distorted; (ii) Fe^{3+} occupies two octahedral sites, one of which is more distorted than the other. From the areas of the fitted doublets, an average $\text{Fe}^{3+}/(\text{Fe}^{2+} + \text{Fe}^{3+})$ value of 0.57(3) is obtained.

From the electron-microprobe data (Table 1), the following chemical formulae are obtained on the basis of 13 oxygen atoms per formula unit (apfu): $(\text{Ca}_{1.17}\text{Ce}_{0.40}\text{La}_{0.19}\text{Nd}_{0.19}\text{Pr}_{0.06})_{\Sigma 2.01}(\text{Fe}^{3+}_{1.43}\text{Fe}^{2+}_{0.71}\text{Al}_{0.61}\text{Ti}_{0.14}\text{Mn}_{0.07})_{\Sigma 2.96}[(\text{Si}_{2.89}\text{Al}_{0.11})_{\Sigma 3.00}\text{O}_{11}]\text{O}(\text{OH})_{1.01}$ for the crystal structurally investigated (column 1, Table 1); $(\text{Ca}_{1.12}\text{Ce}_{0.43}\text{La}_{0.18}\text{Nd}_{0.17}\text{Pr}_{0.05})_{\Sigma 1.96}(\text{Fe}^{3+}_{1.44}\text{Fe}^{2+}_{0.75}$

$\text{Al}_{0.59}\text{Ti}_{0.14}\text{Mn}_{0.12})_{\Sigma 3.04}[(\text{Si}_{2.84}\text{Al}_{0.16})_{\Sigma 3.00}\text{O}_{11}]\text{O}(\text{OH})_{1.01}$ for the mean composition of six different grains (column 2, Table 1); $(\text{Ca}_{1.01}\text{Ce}_{0.50}\text{La}_{0.21}\text{Nd}_{0.18}\text{Pr}_{0.10})_{\Sigma 2.00}(\text{Fe}^{3+}_{1.55}\text{Fe}^{2+}_{0.85}\text{Al}_{0.24}\text{Ti}_{0.20}\text{Mn}_{0.14})_{\Sigma 2.98}[(\text{Si}_{2.89}\text{Al}_{0.11})_{\Sigma 3.00}\text{O}_{11}]\text{O}(\text{OH})_{1.00}$ for the domain poorest in Al (column 3, Table 1).

The Fe_2O_3 and FeO contents were calculated on the basis of the isomorphic substitution $\text{Ca}^{2+} + \text{Fe}^{3+} \leftrightarrow \text{REE}^{3+} + (\text{Fe}, \text{Mn})^{2+}$ and are reasonably confirmed by wet-chemical analysis. Taking into account some zoning and variability in the level of the REE, Ca, Mn and Fe, the partitioning of Fe_{tot} mentioned accumulates the analytical errors on the ratio $\text{Fe}^{3+}/\text{Fe}^{2+}$; the average amount of $\text{Fe}^{3+}/(\text{Fe}^{2+} + \text{Fe}^{3+}) = 0.65$ obtained from the electron-microprobe analysis is consistent with the value of 0.57(3) indicated by the Mössbauer spectra. A simplified formula for ferriallanite-(Ce) is $(\text{Ca}, \text{REE})_2(\text{Fe}^{3+}, \text{Fe}^{2+}, \text{Al})_3(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$ or, ideally, also on the basis of the structure refinement discussed below, $\text{CaCeFe}^{3+}\text{AlFe}^{2+}(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$. The order of the octahedrally coordinated cations is according to the dominance in the sites $M(1)$, $M(2)$ and $M(3)$.

Single-crystal X-ray crystallography

Single-crystal X-ray-diffraction data were collected, from the best of a few selected crystals, on a four-circle Siemens P4 diffractometer and anisotropically refined with the SHELX package (Sheldrick 1997) starting from the coordinates published by Comodi & Zanazzi (1997) for clinzoisite. The relevant experimental data and information on the refinement are given in Table 2. The position of the hydrogen atom was detected in a difference electron-density map and refined. As already found by neutron (Ferraris *et al.* 1989; *cf.* Kvik *et al.* 1988) and X-ray diffraction (Catti *et al.* 1988, 1989) in strontian piemontite, the hydrogen atom is covalently bonded to O(10) [O(10)-H = 0.72(7) Å] and a hydrogen bond is established with O(4) [O(10)...O(4) = 2.889(3) Å]. Table 3 reports the atom coordinates, the anisotropic displacement parameters and the number of refined electrons of the sites *A* and *M*. Table 4 shows the X-ray powder-diffraction data that have been indexed taking into account the calculated intensities of the diffraction maxima. A table with F_o and F_c is deposited at the Depository of Unpublished Data, CISTI, National Research Council of Canada, Ottawa, Ontario K1A 0S2, Canada.

TABLE 1. CHEMICAL COMPOSITION OF FERRIALLANITE-(Ce)

	1		2		3
	wt. %	Range	wt. %	Range	wt. %
SiO ₂	28.72	28.50-30.37	28.09	26.89-30.17	27.52
TiO ₂	1.85	0.93-2.00	1.84	0.73-2.53	2.53
Al ₂ O ₃	6.07	5.41-6.30	6.26	2.83-8.50	2.83
La ₂ O ₃	5.12	3.57-5.44	4.90	2.04-7.14	5.42
Ce ₂ O ₃	10.86	9.96-11.79	11.66	7.69-13.93	13.01
Pr ₂ O ₃	1.63	0.83-1.82	1.46	0.60-2.61	2.61
Nd ₂ O ₃	5.29	4.45-5.36	4.82	3.70-7.93	4.80
CaO	10.85	10.56-10.91	10.33	8.72-14.33	8.98
FeO (total)	25.30	25.02-27.57	25.79	23.29-29.44	27.20
MnO	0.82	0.59-0.94	1.34	0.28-2.13	1.57
Total	96.51		96.49		96.47
Fe ₂ O ₃ *	18.88		18.99		19.62
FeO*	8.44		8.83		9.68
H ₂ O**	1.48		1.48		1.43
Total	100.01		100.00		100.00
Fe ₂ O ₃ + FeO	27.32		27.82		29.30

1. Crystal used for structure refinement, four point-analyses were made with wavelength-dispersion electron microprobe JXA-50A; standards: chkalovite (Si), ilmenite (Ti, Fe), microcline (Al), LaP₂O₇ (La), CeP₂O₇ (Ce), PrP₂O₇ (Pr), NdP₂O₇ (Nd) and wollastonite (Ca).

2. Average of 8 point-analyses from six different grains; ASEM JSM-5300 electron microprobe equipped with energy-dispersion micro-analyzer Link ISIS; standards: chkalovite (Si), metallic Ti, Al, Fe and Mn, nacaphite (Ce), REE(PO₄)₃ for REE.

3. Composition poorest in Al.

* Fe₂O₃ and FeO calculated according to the substitution Ca²⁺ + Fe³⁺ ~ REE³⁺ + $\frac{1}{2}$ (Fe,Mn)²⁺; a wet-microchemical determination gave 4.14% Al₂O₃, 20.09% Fe₂O₃, 8.43% FeO, 1.64% MnO.

** amount of H₂O inferred on the basis of stoichiometry and charge balance; 1.5% was determined by thermal analysis in N₂ atmosphere. No other elements with atomic number higher than 10 were detected by electron-microprobe analysis.

DISCUSSION

The structure

The structure of ferriallanite-(Ce) (Fig. 2) shows the typical features of the monoclinic epidotes, in which two types of chains of edge-sharing octahedra run along [010]. One chain consists of *M*(2) octahedra only; in the second chain the row of *M*(1) octahedra is decorated on both sides by *M*(3) octahedra. The chains are connected both by isolated (SiO₄) and corner-sharing pairs (Si₂O₇) of tetrahedra, thus establishing a three-dimensional framework; the cations at sites *A*(1) and *A*(2) occupy two types of large cavities.

As is the case for the entire epidote group (see references), in particular for allanite-(Ce) (Dollase 1971), the coordination polyhedra for the *A* cations are characterized by a large spread of the cation-oxygen bonds. In Table 5 we report only the *A*-O distances shorter than 3 Å; the relevant effective coordination numbers ECoN (Table 5; Hoppe 1979) show how distorted the coordination polyhedra are, even for these nearest atoms of oxygen. The complete picture is as follows: *A*(1) has a coordination 6 + 1 + 3, having O(9) twice at a distance of 3.168 Å and once at 3.141 Å; *A*(2) has a coordination 6 + 2 + 3 having O(8) twice at 3.053 Å and once at 3.147 Å. The interaction of the cations occupying the two *A* sites with oxygen atoms at a distance of ~3 Å is demonstrated by the charge distribution (CD) calculation (Table 6; see below); the actual second sphere of neighbor oxygen atoms begins only at distances longer than 3.8 Å. Whereas the sites *M*(1) and *M*(2) are only moderately distorted relative to octahedral coordination, *M*(3), as already mentioned, is strongly distorted and shows a 4 + 2 coordination, as occurs also in the other structures of the epidote group, independently of the type of cations (see below).

TABLE 2. MISCELLANEOUS INFORMATION ABOUT THE CRYSTAL AND REFINEMENT DATA FOR FERRIALLANITE-(Ce)

<i>a</i> (Å)	8.962(2)	<i>b</i> (Å)	5.836(2)
<i>c</i> (Å)	10.182(2)	β (°)	115.02(1)
<i>V</i> (Å ³)	482.6(2)		
Space group	<i>P</i> 2 ₁ / <i>m</i>	<i>Z</i>	2
Absorption coefficient (mm ⁻¹)	8.69	<i>F</i> (000)	581.4
Crystal size (mm)	0.07 × 0.18 × 0.18	Radiation	MoK α
2 θ range (°)	70.01	<i>R</i> (int)	0.0355
Collected reflections	4471	Unique reflections	2296
<i>F</i> _o > 4 σ (<i>F</i> _o)	2095		
Method of refinement:		Full-matrix least squares on <i>F</i> ²	
Weights		1/ σ (<i>F</i> _o) ²	
Number of refined parameters	129		
Goodness of fit on <i>F</i> ²	1.124		
<i>R</i> ₁ [<i>F</i> _o > 4 σ (<i>F</i> _o)]	0.026		
<i>wR</i> ₂	0.066		
<i>R</i> ₁ (all data)	0.030		
<i>G</i> _{00F}	1.124		

TABLE 3. FINAL ATOMIC PARAMETERS FOR FERRIALLANITE-(Ce) AND ELECTRONS PER SITE WHERE REFINED (e_{ref})

Site	e_{ref}	x	y	z	U_{00}	U_{11}	U_{22}	U_{33}	U_{23}	U_{13}	U_{12}
A(1)*	21.33(14)	0.75559(8)	3/4	0.484(7)	0.129(2)	0.0199(3)	0.0098(3)	0.0139(3)	0	0.0120(2)	0
A(2)*	53.67(17)	0.5930(2)	3/4	0.42969(2)	0.101(1)	0.0073(1)	0.0141(1)	0.0080(1)	0	0.0023(1)	0
Si(1)		0.3426(1)	3/4	0.03835(9)	0.069(2)	0.0060(3)	0.0072(3)	0.0073(3)	0	0.0027(3)	0
Si(2)		0.6854(1)	1/4	0.27559(9)	0.079(2)	0.0069(3)	0.0083(3)	0.0094(4)	0	0.0043(3)	0
Si(3)		0.1930(1)	3/4	0.32639(9)	0.066(2)	0.0050(3)	0.0075(3)	0.0074(3)	0	0.0027(3)	0
M(1)**	24.08(5)	0	0	0	0.068(1)	0.0057(2)	0.0063(2)	0.0084(2)	0.0005(1)	0.0029(2)	-0.0001(1)
M(2)*	18.73(3)	0	0	1/2	0.077(2)	0.0060(3)	0.0070(3)	0.0093(3)	-0.0007(2)	0.0023(2)	0.0004(2)
M(3)**	25.52(5)	0.30699(6)	1/4	0.21701(5)	0.095(2)	0.0058(2)	0.0106(2)	0.0100(2)	0	0.0012(2)	0
O(1)		0.2382(2)	0.9863(3)	0.0256(2)	0.100(3)	0.0088(6)	0.0068(6)	0.0160(7)	0.0002(5)	0.0068(5)	0.0010(5)
O(2)		0.3169(2)	0.9691(3)	0.3670(2)	0.098(3)	0.0082(6)	0.097(6)	0.0106(6)	0.0003(5)	0.0031(5)	-0.0017(5)
O(3)		0.7955(2)	0.0178(3)	0.3304(2)	0.125(3)	0.0103(7)	0.0077(6)	0.0152(7)	-0.0012(5)	0.0013(6)	0.0004(5)
O(4)		0.0627(3)	1/4	0.1374(2)	0.093(4)	0.0075(9)	0.0107(9)	0.0079(9)	0	0.0016(7)	0
O(5)		0.0525(3)	3/4	0.1558(3)	0.101(4)	0.0098(9)	0.0109(9)	0.0085(9)	0	0.0030(7)	0
O(6)		0.0742(2)	3/4	0.4132(3)	0.104(4)	0.0098(9)	0.0118(9)	0.0130(1)	0	0.0075(8)	0
O(7)		0.5102(3)	3/4	0.1814(3)	0.123(4)	0.0100(1)	0.0150(1)	0.0099(9)	0	0.0020(8)	0
O(8)		0.5445(3)	1/4	0.3325(3)	0.168(5)	0.0080(1)	0.032(1)	0.012(1)	0	0.0055(8)	0
O(9)		0.6041(3)	1/4	0.0977(3)	0.144(4)	0.015(1)	0.020(1)	0.0098(9)	0	0.0064(8)	0
O(10)		0.0937(3)	1/4	0.4316(3)	0.115(4)	0.01(1)	0.014(1)	0.011(1)	0	0.0054(8)	0
H		0.083(8)	1/4	0.358(7)	0.05						

* The scattering curve of Ce was used. ** The scattering curve of Fe was used. * Fe versus Al scattering curve used.

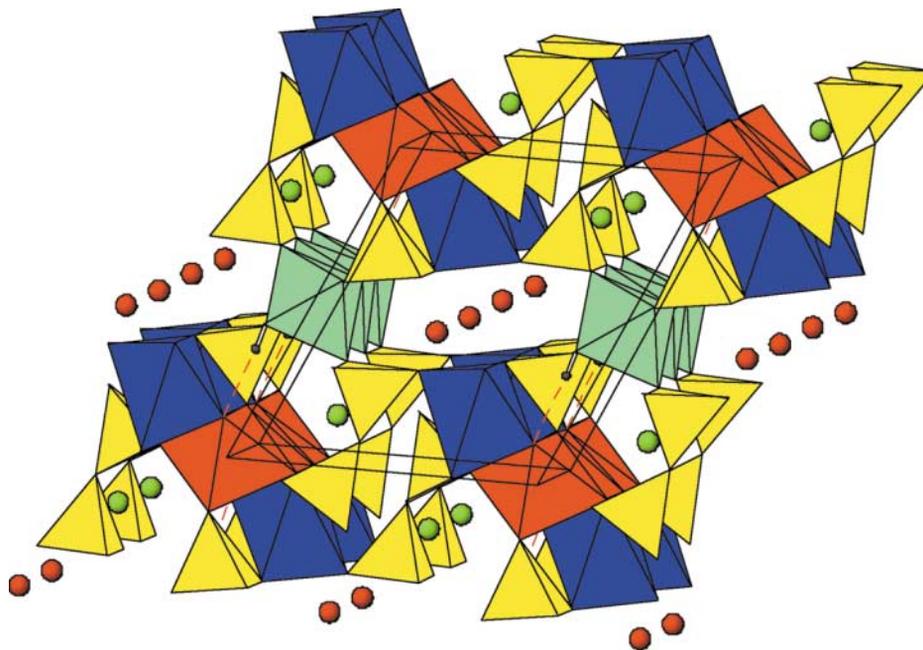


FIG. 2. Crystal structure of ferriallanite-(Ce) seen approximately along [010]. Red, green and blue octahedra represent around the $M(1)$, $M(2)$ and $M(3)$ sites, respectively. The $A(1)$, $A(2)$ sites and the hydrogen atoms are represented by green, red and black circles. The SiO_4 tetrahedra are shown in yellow. The dashed red lines indicate the hydrogen bonds.

TABLE 4. X-RAY POWDER-DIFFRACTION DATA FOR FERRIALLANITE-(Ce)

I_{obs}	I_{calc}	d_{obs}	d_{calc}	hkl	I_{obs}	I_{calc}	d_{obs}	d_{calc}	hkl
20	22	9.30	9.23	001	15	12	2.31	2.301	$\bar{3}04$
40	38	8.07	8.00	101	100	13	2.18	2.224	122
20	17	5.15	5.12	101				2.212	$\bar{1}23$
25	22	5.08	5.03	$\bar{1}02$				2.186	401
15	5	4.92	4.93	011	80	24	2.14	2.170	221
20	20	4.77	4.74	110				2.167	403
15	18	4.59	4.61	002				2.151	$\bar{2}23$
15	11	4.03	4.00	$\bar{2}02$				2.145	014
35	10	3.58	3.62	012	40	16	2.10	2.117	023
55	60	3.55	3.55	$\bar{2}11$				2.067	203
20	16	3.33	3.33	210	15	3	1.96	1.969	$\bar{3}23$
35	14	3.26	3.24	201	20	21	1.90	1.904	$\bar{2}24$
5	4	3.20	3.18	$\bar{2}03$	10	8	1.86	1.888	312
40	31	2.94	2.94	302	10	5	1.82	1.810	024
65	100	2.93	2.93	$\bar{1}13$	10	10	1.75	1.777	422
25	34	2.88	2.92	020	10	10	1.72	1.715	415
20	14	2.85	2.84	211	10	5	1.69	1.695	$\bar{2}06$
80	53	2.72	2.746	$\bar{1}20$	15	16	1.68	1.687	$\bar{1}33$
	40		2.721	013	10	28	1.64	1.650	424
55	21	2.69	2.707	300	35	17	1.63	1.625	$\bar{3}31$
10	6	2.65	2.667	$\bar{3}03$	10	11	1.59	1.592	406
60	61	2.63	2.638	$\bar{3}11$	10	12	1.53	1.563	412
35	34	2.56	2.558	202	50	23	1.46	1.459	040
15	13	2.54	2.524	$\bar{1}22$	15	8	1.42	1.419	215
10	6	2.51	2.516	$\bar{1}04$				1.418	422
15	18	2.42	2.426	$\bar{3}13$	10	3	1.37	1.362	502
55	21	2.34	2.358	222				1.360	026

* CuK α radiation, Guinier camera. Calculated values (I_{calc} and d_{calc}) are taken from the single-crystal data.

TABLE 5. SELECTED INTERATOMIC DISTANCES (Å) AND ANGLES (°) FOR FERRIALLANITE-(Ce)*

$A(1)-O(1) \times 2$	2.370(2)	$M(3)-O(1) \times 2$	2.315(2)
$A(1)-O(3) \times 2$	2.328(2)	$M(3)-O(2) \times 2$	2.217(2)
$A(1)-O(5)$	2.630(3)	$M(3)-O(4)$	1.987(2)
$A(1)-O(6)$	2.979(3)	$M(3)-O(8)$	1.948(3)
$A(1)-O(7)$	2.369(3)	$<M(3)-O>$	2.117
$<A(1)-O>$	2.534		
		$Si(1)-O(1) \times 2$	1.641(2)
$A(2)-O(2) \times 2$	2.492(2)	$Si(1)-O(7)$	1.590(3)
$A(2)-O(2') \times 2$	2.613(2)	$Si(1)-O(9)$	1.646(3)
$A(2)-O(3) \times 2$	2.883(2)	$<Si(1)-O>$	1.626
$A(2)-O(7)$	2.314(3)		
$A(2)-O(10)$	2.555(3)	$Si(2)-O(3) \times 2$	1.630(2)
$<A(2)-O>$	2.571	$Si(2)-O(8)$	1.597(3)
		$Si(2)-O(9)$	1.642(3)
		$<Si(2)-O>$	1.623
$M(1)-O(1) \times 2$	2.040(2)		
$M(1)-O(4) \times 2$	1.933(2)	$Si(3)-O(2) \times 2$	1.628(2)
$M(1)-O(5) \times 2$	2.058(2)	$Si(3)-O(5)$	1.659(3)
$<M(1)-O>$	2.010	$Si(3)-O(6)$	1.646(2)
		$<Si(3)-O>$	1.644
$M(2)-O(3) \times 2$	1.917(2)		
$M(2)-O(6) \times 2$	1.962(2)	$Si(1)-O(9)-Si(2)$	141.0(2)
$M(2)-O(10) \times 2$	1.953(2)		
$<M(2)-O>$	1.964		

* See text for additional distances at the sites A.

Crystal chemistry

The following features are crucial for the determination of the crystal-chemical formula of ferriallanite-(Ce).

TABLE 6. CHARGE AT THE SITE (C_{obs})¹ COMPARED WITH THE CALCULATED CHARGE-DISTRIBUTION (CD) AND BOND-VALENCE (BV) FOR FERRIALLANITE-(Ce)

Site	C_{obs}	CD	BV	ECoN*	Site	C_{obs}	CD	BV
A(1)	2.03	2.03	2.13	5.91	O(1)	-2	-1.94	-1.94
A(2)	2.89	2.78	2.83	7.49	O(2)	-2	-2.07	-1.93
M(1)	3.14	3.19	3.07	5.83	O(3)	-2	-2.04	-2.04
M(2)	3	2.93	3.15	5.98	O(4) [‡]	-2	-2.00	-1.78*
M(3)	2	2.04	2.09	4.71	O(5)	-2	-1.96	-1.97
Si(1)	4	4.06	3.93	3.97	O(6)	-2	-1.93	-2.03
Si(2)	4	4.12	3.98	3.99	O(7)	-2	-2.10	-2.02
Si(3)	3.93	3.93	3.81	3.99	O(8)	-2	-1.84	-1.94
H	1	0.93	0.98	1.45	O(9)	-2	-1.89	-2.09
					OH**	-1	-2.20	-1.35*

¹ According to the assigned site-population (Table 7). * Effective coordination-number for cations (see text). ** OH = O(10), donor of the hydrogen bond, the neutron coordinates of Ferraris *et al.* (1989) were used for the calculation of CD (see text). ‡ Acceptor of the hydrogen bond. † Without the contribution of the H atom.

TABLE 7. SITE OCCUPANCY FOR FERRIALLANITE-(Ce) ON THE BASIS OF THE NUMBER OF REFINED ELECTRONS (e_{ref}) PER SITE

Site	e_{ref}	Site occupancy ^a	* e_{top}	Charge	* L_{top}	L_{obs}
A(1)	21.33(14)	$Ca_{0.97}Ce_{0.03}$		2.03	2.521	2.534
A(2)	53.67(17)	$Ce_{0.89}Ca_{0.11}$	71.42	2.89	2.540	2.571
M(1)	24.08(5)	$Fe^{3+}_{0.80}Ti^{4+}_{0.14}Al_{0.06}$		3.14	1.957	2.010
M(2)	18.73(3)	$Al_{0.36}Fe^{3+}_{0.44}$		3.00	1.942	1.964
M(3)	25.32(5)	$Fe^{2+}_{0.93}Mn_{0.07}$	69.32	2.00	2.014	2.117
Si(3)	14 (fixed)	$Si_{0.94}Al_{0.06}$	13.89	3.94	1.668	1.644

^a The REE site is assumed to be populated by Ce only for purpose of site-occupancy refinement. * The number of electrons expected from the chemical analysis of the crystal used for the structure refinement (see text); cumulative values normalized to full occupancy are given for the sites A and M. † Expected site-oxygen bond lengths (Å) for the compositions shown according to the following ionic radii (Shannon 1976): O 1.40, Ca 1.12, Ce 1.143, Fe³⁺ 0.55, Fe²⁺ 0.61, Mn 0.67, ²⁷Al 0.535, ²⁹Al 0.39, Si 0.26 Å. ‡ Observed average bond-lengths (Å) (Table 3).

1. The chemical composition (Table 1); in particular, the total amount of Fe must be distributed over three M sites.

2. For the refined number of electrons at the cation sites where mixed occupancy occurs (Tables 3, 7), most of the Al must stay in the M(2) site, *i.e.*, the poorest in electrons, as usually found for epidotes, in particular for allanite-(Ce) (Dollase 1971).

3. According to the results of Mössbauer spectroscopy (see also Dollase 1973), Fe²⁺ occurs in only one octahedral site, which is distorted, and the amount of Fe³⁺ is split between two octahedral sites. Neither of these two sites can coincide with that assigned to Fe²⁺ (Table 7).

4. According to the values of the effective coordination-number (ECoN, Table 6; Hoppe 1979), M(3) is the most distorted octahedral site. This is a constant feature of the epidote-group structures and is related to the

absence of symmetry constraints on $M(3)$, as discussed by Ferraris *et al.* (1989).

5. The small amount of tetrahedrally coordinated Al should reasonably be located in the larger (Table 5) and isolated $\text{Si}(3)\text{O}_4$ tetrahedron.

On the basis of the features listed above, the following crystal-chemical formula is assigned to the crystal of ferriallanite-(Ce) whose structure has been refined (all REE are ideally indicated as Ce): $(\text{Ca}_{0.97}\text{Ce}_{0.03})(\text{Ce}_{0.89}\text{Ca}_{0.11})(\text{Fe}^{3+}_{0.80}\text{Ti}^{4+}_{0.14}\text{Al}_{0.06})(\text{Al}_{0.56}\text{Fe}^{3+}_{0.44})(\text{Fe}^{2+}_{0.93}\text{Mn}_{0.07})(\text{Si}_{0.94}\text{Al}_{0.06}\text{O}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$. The ratio Ca/Ce , 1.17 in this formula, is constrained by the total number of electrons (75.00, Table 7) belonging to the two A sites according to the structure refinement. The value of this ratio differs from 1.39, obtained by chemical analysis. Owing to charge balance, this difference is reflected in the ratio $\text{Fe}^{3+}/(\text{Fe}^{2+} + \text{Fe}^{3+})$, which is 0.57 in the structural crystal-chemical formula, instead of 0.65 as in the analytical formula given above. However, the former value corresponds exactly to that indicated by the Mössbauer analysis. Reasonably, the difference is due to zoning and analytical errors, mainly involving REE, as mentioned above.

The partitioning of cations shown in the structural crystal-chemical formula has been checked by calculations of charge distribution (CD; Hoppe *et al.* 1989, Nespolo *et al.* 1999, 2001) and bond valence (BV; Brown & Wu 1976) in light of competing hypotheses, like the assignment of Fe^{2+} to the site $M(1)$ (Table 6). For a comparison with the values of the mean bond-lengths (Table 5), Table 7 also gives the values of the bond lengths to be expected for the mixed cations indicated according to the ionic radii published by Shannon (1976). To complete the number of positive charges in the structure, the position of the H atom obtained in the neutron-diffraction study of a sample of strontian piemontite by Ferraris *et al.* (1989) has been used for the CD calculation. In fact, the position obtained in the present study is too approximate, which is not unusual for X-ray-diffraction data. Of course, the position of H cannot be exactly the same in the two structures, and that may explain the high charge obtained for the OH group, O(10).

The assignment of Fe^{2+} to $M(3)$ is in accordance with the results published by Dollase (1973) for allanite-(Ce), where he also estimated a concentration of Fe^{3+} higher in $M(1)$ than in $M(3)$, *i.e.*, the reverse of the normal distribution in epidote-group minerals. Actually, as mentioned above, in ferriallanite-(Ce) the concurrent high occurrence of Fe^{2+} , which occupies only one site, and Fe^{3+} , which partitions between two sites, excludes the presence of Fe^{3+} in the same site as Fe^{2+} .

CONCLUSIONS

The majority of allanite-subgroup minerals contain both Fe^{2+} and Fe^{3+} ; in general, these minerals have a composition $(\text{Ca}_{2-x}\text{REE}_x)(\text{Al}_{3-x-y}\text{Fe}^{2+}_x\text{Fe}^{3+}_y)(\text{SiO}_4)$

$(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$, where $0.50 < x \leq 1.00$ and $x + y < 2.49$. A preference of y around 0.3 is noted (Kepezhinskas & Khlestov 1971); the highest value of y , 0.84, is reported by Krivdik *et al.* (1989). The present description of ferriallanite-(Ce) shows that members with $y > 1$ and Fe^{3+} dominant in the site $M(1)$ occur, whereas Al still dominates in $M(2)$. In an allanite-(Ce) – ferriallanite-(Ce) solid solution, ideally the end members have formula $\text{CaCeFe}^{2+}\text{Al}_2(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$ [allanite-(Ce)] and $\text{CaCeFe}^{3+}\text{AlFe}^{2+}(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$ [ferriallanite-(Ce)]. The composition poorest in Al reported in Table 1 corresponds to a REE-bearing member of the epidote group with ideal formula $\text{CaCeFe}^{2+}\text{Fe}^{3+}_2(\text{SiO}_4)(\text{Si}_2\text{O}_7)\text{O}(\text{OH})$. However, insufficient material is available for a full characterization of this potentially new epidote.

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